You have cut and pasted the element blocks and ARROWS at this point and you are ready to answer the questions!!!!!!

Periodic table basics:

1. **What happens to the size of the atomic radius of an atom going down a group?** increase in size
2. **What happens to the size of the atomic radius of an atom going across a period?** decrease in size
3. **A** ion **is when an atom or group of atoms have gained or lost an electron.**
4. **The energy required to remove one electron from a neutral atom of an element is the** ionization energy.
5. **The green colored elements represent the** alkali **family.**
6. **The purple colored elements represent the** halogen **family.**
7. **The blue colored elements represent the** alkaline earth **family.**
8. **The yellow colored elements represent the** noble gas **family.**
9. **Which elements have full s and p orbitals?** noble gases He, Ne, Ar
10. **Which family only has one valance electron?** alkali
11. **Predict the number of valance electrons for each element based on its location in the periodic table of elements. You will need to use your table or textbook.**

Barium= 2 **Lead=** 2 **Xenon=** 8 **Potassium=** 1

12. **Why is hydrogen in group IA when it is a non-metal?**
Has 1 valance e⁻; loses 1 valance e⁻; has +1 charge
13. **Write the electron configuration for silver, Ag** ← this is an exception!
$5^2d^9 \rightarrow 5^1d^{10}$