

Name: _____

Period: _____

Ionic/Covalent/Acids Nomenclature PowerPoint Questions

1. What is the difference between a binary and ternary compound?
2. What is the difference between a monatomic and polyatomic ion?
3. Write a brief statement of how to write the **name** of a binary ionic compound.
4. Write a brief statement of how to write the **name** of a binary ionic compound that contains a metal with variable oxidation numbers (charges).
5. Read over all the symbol examples with their charges.
6. Look at the examples of how to use the criss cross method to balance ionic compound formulas. What was the formula for Strontium Fluoride? Copper(I) Sulfide? And Lead(IV) Oxide? What did you need to do with the subscripts of Lead(IV) Oxide to get the final answer?
7. Read over the polyatomic ion list. Have you memorized them yet?
8. Look at the ternary ionic compound examples. When do you need to use parentheses?
9. What is the TOTAL number of atoms in the compound $\text{Al}_2(\text{SO}_4)_3$?
10. Read over all the examples of ternary ionic compounds being balanced using the criss cross method. What was the formula for calcium phosphate? Ammonium carbonate? Copper(II) hydroxide?
11. Read over the lists of ionic compound examples. Pay attention to when there is a roman numeral or parentheses. What did you get on the POP QUIZ on numbers 2, 5, 7 and 9?
12. What type of elements are covalent compounds composed? Are oxidation numbers (charges) needed? How do you use the prefixes to name a covalent compound?
13. What is the prefix for four? Seven?
14. What is the formula for diarsenic trisulfide? What is the name of N_2O_5 ?
15. Acid names are derived from the anion that is bonded to hydrogen. -ide ending names change their ending to _____? -ite ending names change their ending to _____? -ate ending names change their ending to _____?
16. Read over all the examples of the acids. What do all of these acids start with? If an acid does NOT contain oxygen and only has a ONE-element anion, what prefix is included in its name?