

Name _____ Period _____ Date _____

Naming Binary Ionic Compounds

METAL WITH A NONMETAL ALSO CALLED BINARY COMPOUNDS.

1. The name of the metal is written first. The metals are in column IA, IIA, aluminum, zinc, silver and cadmium.
2. The ending of the nonmetal is changed to - IDE.

EXAMPLE: NaCl is **sodium chloride**.

K₂O is **potassium oxide**.

Be₃N₂ is **beryllium nitride**

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| _____ 1. NaBr | _____ 2. KI |
| _____ 3. SrCl ₂ | _____ 4. CaI ₂ |
| _____ 5. Mg ₃ N ₂ | _____ 6. Li ₃ P |
| _____ 7. BaI ₂ | _____ 8. NaF |
| _____ 9. Al ₂ O ₃ | _____ 10. K ₂ S |
| _____ 11. CdBr ₂ | _____ 12. LiCl |
| _____ 13. Ag ₃ N | _____ 14. Al ₂ S ₃ |
| _____ 15. MgF ₂ | _____ 16. CdI ₂ |
| _____ 17. CaO | _____ 18. CsBr |
| _____ 19. Ag ₂ S | _____ 20. Na ₂ O |
| _____ 21. Ba ₃ N ₂ | _____ 22. SrCl ₂ |
| _____ 23. LiF | _____ 24. CdS |
| _____ 25. AlP | _____ 26. MgF ₂ |
| _____ 27. BeS | _____ 28. NaI |
| _____ 29. Ag ₂ O | _____ 30. K ₂ O |
| _____ 31. CdCl ₂ | _____ 32. AgF |
| _____ 33. Rb ₃ N | _____ 34. Sr ₃ N ₂ |
| _____ 35. ZnF ₂ | _____ 36. CdS |
| _____ 37. CsCl | _____ 38. ZnBr ₂ |
| _____ 39. KF | _____ 40. AlF ₃ |

Endings of some common elements changed to -ide.

Fluorine	F ⁻	Fluoride	Sulfur	S ²⁻	Sulfide
Chlorine	Cl ¹⁻	Chloride	Nitrogen	N ³⁻	Nitride
Bromine	Br ¹⁻	Bromide	Phosphorus	P ³⁻	Phosphide
Iodine	I ¹⁻	Iodide	Carbon	C	Carbide
Oxygen	O ²⁻	Oxide	Hydrogen	H ¹⁺	Hydride