Name _____ Period _____ Date_____

Naming Binary Ionic Compounds containing Transition Metals

To write the formula for an ionic compound containing a metal with variable charge requiring a **Roman Numeral.** The Roman Numeral indicates the positive charge of the metal. Change the ending of the nonmetal to **ide**.

Roman Numerals

One is I. Two is II. Three is III. Four is IV. Five is V.
Six is VI. Seven is VII. Eight is VIII. Nine is IX.

1. $FeCl_3$ _____ **2. $FeCl_2$**_____

3. Fe_2O_3 _____ **4. CuO** _____

5. CuBr _____ **6. $PbCl_4$** _____

7. $PbCl_2$ _____ **8. PbO** _____

9. $SnCl_4$_____ **10. $SnCl_2$** _____

11. Cu_2S_____ **12. FeI_3** _____

13. PbO_2_____ **14. Cu_2O** _____

15. Cu_3N_2 _____ **16. FeI_2**_____

17. $CuCl_2$ _____ **18. FeS**_____

19. Cu_3P _____ **20. Sn_3N_2** _____

21. Cu_3N_____ **22. PbS_2**_____

23. CrI_3 _____ **24. PbI_2** _____

25. SnF_4 _____ **26. $SnCl_2$**_____

27. SnS _____ **28. SnO** _____

29. Fe_2S_3 _____ **30. Fe_3N_2** _____

31. SnS_2_____ **32. SnO_2**_____

33. $NiCl_2$_____ **34. NiO**_____

35. Co_2S_3_____ **36. CoP**_____

Endings of some common elements changed to -ide.

Fluorine		**F^{1-}**	**Fluoride**	**Sulfur**		**S^{2-}**	**Sulfide**
Chlorine		**Cl^{1-}**	**Chloride**	**Nitrogen**		**N^{3-}**	**Nitride**
Bromine		**Br^{1-}**	**Bromide**	**Phosphorus**		**P^{3-}**	**Phosphide**
Iodine		**I^{1-}**	**Iodide**	**Carbon**		**C**	**Carbide**
Oxygen	**O^{2-}**	**Oxide**		**Hydrogen**	**H^{1+}**	**Hydride**	