

Name _____ Date _____ Period _____

Naming Molecular Compound

MOLECULAR COMPOUNDS- COMPOSED OF NONMETALS ONLY.

1. Name of the "more metallic" element is written first. Since both are nonmetals, choose the element closest to the bottom or left of the periodic table.
2. The ending of the second nonmetal is changed to -ide.
3. Prefixes are added to indicate the number of atoms present. Mono is not used on the first element just write the name if there is only one.

1	MONO	4	TETRA	7	HEPTA		
2	DI	5	PENTA	8	OCTA		
3	TRI	6	HEXA	9	NONA	10	DECA

EXAMPLES:

CO is **carbon monoxide**

CO₂ is **carbon dioxide**

SF₆ is **sulfur hexafluoride**

P₂O₃ is **diphosphorus trioxide**

Endings of some common elements changed to -ide.

Fluorine F ¹⁻	Fluoride	Chlorine	Cl ¹⁻	Chloride
Bromine	Br ¹⁻ Bromide	Iodine	I ¹⁻	Iodide
Oxygen O ²⁻	Oxide	Sulfur	S ²⁻	Sulfide
Nitrogen N ³⁻	Nitride	Phosphorus	P ³⁻	Phosphide
Carbon C ^{+or-4}	Carbide	Hydrogen	H ¹⁺	Hydride

Write the correct name for the following compounds.

1. KrF₂ _____
2. BrCl₅ _____
3. SCl₄ _____
4. PF₃ _____
5. CO _____
6. PCl₅ _____
7. BrO₃ _____
8. BN _____
9. N₂S₅ _____
10. NI₃ _____
11. SF₆ _____
12. XeF₄ _____
13. P₂O₅ _____
14. S₂Cl₂ _____
15. As₄O₁₀ _____